



High School Science Virtual Learning

College Chemistry

Titration Virtual Lab

May 12, 2020



High School College Chemistry

Lesson: May 12, 2020

Objective/Learning Target:

Students will check their work from the previous lesson, and watch a video to expand their knowledge of titrations.



Let's Get Started:

1. What is the purpose of a titration?
2. Which type of solutions are commonly used in titrations?



Let's Get Started: **Answer Key**

1. What is the purpose of a titration? **Titration**s are used to **determine the concentration of an unknown solution.**
2. Which type of solutions are commonly used in titrations?
Acids and bases



Lesson Activity:

Directions:

1. Check your answers to your worksheet from the previous lesson with this [answer key](#).
2. Watch this Teacher's Pet [video](#) about titrations.



Practice

Complete the following questions using the information you learned during the lesson activity.

Questions:

1. What are some solutions that might need to be titrated for something other than acids or bases?
2. Why is it important to add the indicator to the titration?
3. In a titration, 16.08 mL of NaOH react with 0.0052 moles of KHP. What is the molarity of the base?

Once you have completed the practice questions check with the **answer** key.

1. Any drink that contains salt could be titrated, such as energy drinks.
2. The indicator tells you when the reaction is complete. Without it, you would titrate endlessly.
3. $M = \text{mol/L} = 0.0052 \text{ mol} / 0.01608 \text{ L} = 0.32 \text{ M}$
(The moles of KHP is equal to moles of NaOH)



Additional Practice:

Click on the link below to see a titration that is not used for acids or bases. This titration can be used as part of the process to determine the concentration of bleach.

[Thiosulfate Titration Video](#)